**Q1/** Calculate the value of Relative Atomic Mass **(*Ar)*** for naturally occurring chlorine and ratio of **(35Cl:37Cl)** if the distribution of isotopes is **75.77% (1735Cl )** and **24.23% (1737Cl)** . Accurate mass for 35Cl and 37Cl are **34.97** and **36.97** respectively.

**Note: The relative atomic Mass (Ar) = mean No. of the two isotopes.**

**Ans.//**



**Q2/** Element **(X)** is toxic to human in high concentration, but is essential to life at low concentration. Identify (X) element whose atoms contain **24-Protons** and the isotope with **28-Neutrons** write the symbol of element **(X)**?

Ans.//

 Atomic No. (Z)= No. of proton = No. of electron = 24

No. of neutron = 28

A= Z+N No. of mass (A) =24+28= 54 

 X is (Cr) Chromium element