Salahaddin University-Erbil

College of Science

Department of Biology

1st Year Students
By Nishtiman S hasan

Subject: Practical General Chemistry

Q1/ Describe the following :-

Sublimation , freezing point

Q2/ Define Boiling point and melting point

Q3/ A/How are impurities removed during re-crystallization? Write properties of the best solvent for recrystallization of solid compound. Q4/ Write the main factors can be influence melting point determination.

Q5/ The freezing point of (8ml) of dioxine is equal to (–0.5C0), when we put (0.4gm) of camphor the freezing point will be decreased to(-4C0), calculate M.wt of the camphor. Kf =5.12

Q6/ Give the reason

1. Using chloroform to extract caffeine.

Q7/ Why Using water to extract caffeine from tea leaves.

Q8/ Define the following :-

1. Standard solution

2. Boiling point

Q9/ Use hot water to extract the caffeine from tea leaves.

Q10/ Not use NaoH as the primary standard.

Q11/ How are impurities removed during Re-crystallization?

Q12/ The freezing point of (8ml) of dioxine is equal to (–0.5C0), when we put (0.4gm) of camphor the freezing point will be decreased to(-4C0), calculate M.wt of the camphor. Kf =5.12

Q13/ Write the main differences between End point and Equivalence point ?

Q14/ Write a short report about precipitation reaction.

Q 15/ Write the types of hardness, the methods of releasing them.

Q16 / Describe the followings:

Purification melting point

Q17/ Write two methods for purification of solid materials.

Q18/ why does diet soda freeze at a higher temperature that regular soda?

Q19/ In order to find the molar mass of an unknown compound, a research scientist prepared a solution of 0.930 g of unknown in 125 g of a solvent. The pure solvent had a freezing point of 74.2 °C, and the solution had a freezing point of 73.4 °C. Given the solvent's freezing-point depression constant, $Kf = 5.50$ °C/m, find the molar mass of the unknown.

Q20/ Write aim of re-crystallization.