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University of Salahaddin-Hawler

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Practical Representative Elements, Second Stage

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Week 1

• Error

Classes of Errors







Error refer to the **difference** between a **measured** value and **true** value



Classes of Errors

- 1-Determinate Errors (Systematic Errors)
- A-Instrumental Errors
- B-Operative Errors
- C-Personal Errors
- D-Methodic Errors (Errors of Method)

2-Indeterminate Errors (Random Errors)



Classes of Errors

• 1-Determinate Errors (Systematic Errors)

Are these errors that are **known** and **controllable** errors (are **known, avoided and corrected**).

Determinate errors are those whose magnitude can be determined (at least in principle) and the measurements thereby corrected. **The analyst can measure and account for these errors**.

- A- Instrumental Errors: Those due to apparatus and reagents
- **1-Balance and Weight;**
- uncalibrated weight



2- Volumetric apparatus: Use of uncalibrated glassware.



Table 4.2	Measurement Errors for Type A Volumetric Glassware [†]				
Transfer Pipets		Volumetric Flasks		Burets	
Capacity (mL)	Tolerance (mL)	Capacity (mL)	Tolerance (mL)	Capacity (mL)	Tolerance (mL)
1	± 0.006	5	± 0.02	10	± 0.02
2	± 0.006	10	± 0.02	25	± 0.03
5	± 0.01	25	± 0.03	50	± 0.05
10	± 0.02	50	± 0.05		
20	± 0.03	100	± 0.08		
25	± 0.03	250	±0.12		
50	± 0.05	500	± 0.20		
100	± 0.08	1000	± 0.30		
		2000	± 0.50		

- A- Instrumental Errors: Those due to apparatus and reagents
- 3-Vessels and utensils; loss in weight of platinum

crucibles when strongly heated.



4- Reagents; Presence of impurities or interfering substances.

• B- Operative Errors:

These errors are **mostly physical in nature** and are associated with the manipulations of an analysis, and their **magnitude depends more upon** the **analyst** himself than on any other factor.

These errors are <u>reduced</u> to in significance by **Careful, skillful**, and **understanding work**.

- C- Personal Errors:
- 1- Some **persons** are unable to **judge color** changes exactly.
- 2- Taken in **reading a scale**, burette reading.



- D- Errors of Methods: Methodic Errors
- Some sources of methodic errors are these:
- 1- Solubility of a precipitate in the solution in which it is precipitated, and in the wash liquid.

2- Failure of a reaction to proceed to quantitative competition.

3-Precipitation of another substance with the reagent used.

- D- Errors of Methods: Methodic Errors Some sources of methodic errors are these:
- 4- Coprecipitation and Post Precipitatioin.

5-Decomposition or volatilization of a precipitate on heating.

6-Induced reaction and side reactions.

2-Indeterminate Errors: (Random Statistical)

 It is generally stated that random errors are beyond the power of the observer to prevent or allow for, but this statement refers to unchanging conditions and methods of measurement. By working under same conditions and by taking certain precautions.

Homework



• Define The following terms:

1-Salt

- 2-Precipitant Agent
- **3-Decomposition reaction**
- **4-Coprecipitation**
- **5-Post precipitation**
- **6-Solubility**





Question

Write the **chemical formula** for the following chemicals and then calculate their **molecular weights**?

- 1-Water
- 2-Sodium chloride
- 3-Hydrochloric acid
- 4-Barium chloride
- 5-Carbon dioxide

• <u>At. Wt.:</u>

H:1, O:16, Ba:137.32, Cl:35.45, Na:23, C: 12