Department of Chemistry

College of Science

University of Salahaddin

Subject: Electro-Chemistry

Course Book – *third stage*

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Assistant: *Hawtta .* (for Practical)

Experiment -1-Determination of cell constant.

Experiment -2-Dissociation constant of weak acid from conductivity measurement.

Experiment -3-Determination of equivalent conductance of strong electrolytes.

Experiment -4-Conduct metric titration.

a) Strong acid x Strong base.

b) Weak acid x Strong base.

c) Mixture of acids X Strong base.

d) Weak acid x Weak base

e) Precipitation titration.

F) Displacement titration.

Experiment -5-Determination of solubility product of sparingly soluble salt by conduct metric method.

Experiment -6-Determination the degree of hydrolysis and hydrolysis constant of salt by conduct metric method.

Experiment -7-Effect of viscosity on conductance of solutions.

Experiment -8-Thermodynamic constants for chemical cell.

Experiment -9-a-Application of Nernest equation to Cu/Cu+2 electrodes.

Experiment -9-b-Application of Nernest equation to Zn/Zn+2 electrodes.

Experiment -10-Determination of electrode Potential.

Experiment -11-Potentiometric determination of the dissociation constant of weak acid.

Experiment -12-Determination of decomposition Potential:

Experiment -13-Determination of Avogadro numbers and Faradayconstant from electrolysis of dilute sulphuric acid.